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(54) A process for the preparation of supported metallocene catalyst

(57) This invention relates to a process for the preparation of supported metallocene catalyst, which comprises preparing a precursor catalyst material which consists of preparing the solution of magnesium halide compound in an electron donor solvent in which the magnesium halide being completely soluble, heating the solution to a temperature in the range of 65°C to the boiling point of the respective electron donor for a period ranging between 10 to 60 minutes, separately preparing a solution of the metallocene compound into the same electron donor solvent, heating the solution to a temperature in the range of 25° C to 70° C for a period ranging between 0.1 to 0.5 hrs., mixing the two solutions to obtain the homogenous solution of catalyst precursor compound, stirring and maintaining this resulting homogenous solution at a temperature in the range of 50° C to 70° C for a period of 0.5 to 2 hrs., separately preparing the slurry of the support in the liquid hydrocarbon medium, heating the slurry to a temperature in the range of 50° C to 75° C and stirring it for a period of 0.25 to 2 hrs., mixing the homogenous solution of the catalyst precursor compound with the slurry of the support, stirring the slurry for 0.5 to 3 hrs maintaining at a temperature ranging between 50° C to 70° C, cooling the resultant slurry to room temperature under inert atmosphere, separating the solid by conventional methods, washing the solid by hydrocarbon solvent, drying the solid under vacuum to obtain the product.

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Description

FIELD OF THE INVENTION

This invention relates to a process for the preparation of supported metallocene catalyst. More particularly it relates to the process for the preparation of the said catalyst which is supported with silica-magnesium halide. The catalyst prepared by the process of the present invention is capable of olefin polymerization and copolymerization, giving olefin polymers and copolymers with high catalyst activity at high temperatures.

BACKGROUND OF INVENTION

Anhydrous magnesium chloride, commonly used as a support in high activity olefin polymerization catalyst is, often not, convenient because it is very brittle and undergoes easy attrition in the polymerization reactor. Silica is a well known support for fluidized bed gas phase polymerization of olefins using titanium based Ziegler-Natta catalysts. However, when used with metallocene type catalysts silica supports show low polymerization activities. Therefore, a combination of the two, namely, silica and anhydrous magnesium chloride offers a good balance of properties to prepare catalysts with high activity controlled particle size and good resistance to attrition.

Olefin polymerization and copolymerization catalysts comprising of a metallocene and an aluminum alkyl component are well known in the prior art. The molecular weight of the polymer can be controlled by adjusting the reaction temperature or the amount of co-catalyst or by the addition of hydrogen. The metallocene catalysts require the use of methylaluminoxane (herein after referred to as MAO) as co-catalyst, which is produced by reacting an aluminum alkyl with water.

PRIOR ART REFERENCE

The principle disadvantages of these soluble homogeneous metallocene-methylaluminoxane catalyst system are the need for a large molar excess of MAO), (Al/Metal > 105) for obtaining acceptable polymerization activities and poor control on polymer powder morphologies. Furthermore, when used in gas or slurry phase processes there is a tendency for reactor fouling by forming deposits on the surface of the walls of the reactor and stiffer. These deposits result in the agglomeration of the polymer particles when the metallocene and aluminoxanes or both exists in the suspension medium. Such deposits in the reactor system must be removed regularly. Otherwise, they prevent adequate heat removal from the reaction, adversely affecting the product quality.

The above disadvantages can be obviated by the use of heterogeneous catalysts. Heterogeneous metallocene based catalysts are generally prepared using

either a porous inorganic oxide carrier such as silica, alumina, magnesium oxide and silica/alumina or a group IIA alkaline earth metal halides, preferably, anhydrous magnesium chloride. Attempts have been made to support the catalyst components either on a refractory oxide or by mixing methylaluminoxane and metallocene with inorganic supports (Eur. Pat. Appl. No. 206,794). Eur. Pat. Nos. EP 285, 443 EP 295, 312 describe a polymerization system, in which, an inorganic oxide and an organic polymer are employed as support for metallocenes. The cocatalyst, MAO, by the use of solvent in which it is poorly soluble, is precipitated on the support. JP 06, 184, 226 describes a silica supported metallocene catalyst, which, in conjunction with MAO is useful for the polymerization of ethylene with moderate activity. Eur. Pat. No. EP 619, 325 describes a silica supported metallocene catalyst which lead to the production of bimodal polyethylene. Eur. Pat. No. EP 619, 326 discloses another silica supported metallocene catalyst which can be used for gas-phase polymerization of ethylene in combination with N,N-dimethylanilinium (tetrakis(pentafluorophenyl)borate and triisobutylaluminum as cocatalyst.

As an alternative, a catalyst has been prepared by reacting water containing silica with trimethylaluminum and to use the resultant suspension as a support for metallocenes (Eur. Pat. No. 323, 716) or to carry out the entire reaction of trimethylaluminum, water containing silica and metallocene in the polymerization reactor (Eur. Pat. No. 170, 059). These processes lead to the formation of MAO dissolved in the reaction medium which eliminates the advantages of the support. Furthermore, these supported methylaluminoxane have significantly lower polymerization activity.

A number of patents describe silica supported zirconium based metallocene catalyst systems (US. Pat. No. 5,057, 475; WO 93/09148; JP 05, 170, 822; JP 06, 184, 226; JP 07, 53, 624). Eur. Pat. No. 608, 054 and WO 94/21691 describe highly active silica supported metallocene catalysts which in combination with MAO or trialkylaluminum compounds can be used for homo or co-polymerization of olefins.

Eur. Pat. No. 435, 514 and 436, 326 discloses a supported zirconium based metallocene catalysts prepared by reaction of dibutylmagnesium and tertiary butyl chloride in presence of an electron donor such as diisobutyl ether. These catalysts are useful for the gas-phase polymerization of ethylene. JP 03, 66, 710 and JP 05, 186, 524 describes a magnesium chloride supported metallocene catalysts which is useful for propylene polymerization.

However, both silica and magnesium chloride when individually used as supports have certain drawbacks. Magnesium chloride is brittle in nature. It has a tendency to undergo attrition during its use as the catalyst is subjected to high exothermicity. These catalysts, therefore, cannot be employed at high temperatures. On the other hand, silica, when used as a support with met-

allocene type catalysts show low polymerization activities and, often, causes a broadening of molecular weight distribution. Therefore, it may be advantageous to use a combination of both silica and magnesium chloride, to confer up on the catalyst, the specific advantage of each of the supports.

US Pat. No. 5, 032, 562, describes the preparation of a solid supported catalyst, by the reaction of dibutyl magnesium; a zirconium based neutral metallocene and a compound of a transition metal halide such as titanium tetrachloride and impregnating the said precursor onto silica in the presence of an aluminosilicate activator. However, this catalyst results in bimodal distribution of molecular weights, which is not desirable in most of the applications where polyethylene is used.

JP 04, 96, 908, describes another supported solid catalyst prepared by reacting aqueous magnesium halides in presence of silica, metallocene and an organoaluminum activator, which shows only moderate activity towards ethylene polymerization at 80°C.

Eur. Pat. No. EP 613, 908, describes a silica-magnesium chloride supported metallocene catalyst which in the presence of an organoboron compound polymerizes ethylene with high yield. However, the catalyst prepared according to this method is not very stable to storage and handling.

SUMMARY OF THE INVENTION

Therefore, the present invention is to provide an olefin polymerization and copolymerization catalyst capable of producing high catalyst activity, narrow molecular weight distributions, especially at high temperatures and capable of being used either in gas or slurry phase processes, and in addition, to provide a catalyst by the deposition of a homogeneous solution of anhydrous magnesium chloride and a zirconium based metallocene on to a carrier like high surface area silica, affording a superior impregnation of magnesium-zirconium catalyst precursor on the support.

DETAILED DESCRIPTION OF THE INVENTION

The solid catalyst for olefin polymerization and copolymerization according to the present invention is characterized in that the said catalyst comprises of:

- a) a particulate carrier composed of an oxide of at least one selected from among those belonging to the groups II, III, IV of the Periodic Table having a surface area of about 300 m² /g and containing at least 3 mmol of hydroxyl group per gram of the oxide.
- b) a transition metal compound of a metal belonging to the group IVB of the Periodic Table containing a ligand having a cyclopentadienyl skeleton.
- c) a particulate compound magnesium halide has the structure MgX₂ where, X is selected from the

group containing Cl, Br, or I, wherein the magnesium halide and the transition metal compound are supported on the particulate carrier.

An object of the present invention is to provide a process for the preparation of supported metallocene catalyst designed to rectify the drawbacks in the prior art catalyst by furnishing a new solid supported catalytic system, whose preparation can be performed easily.

Another object of the present invention is to provide a process for the preparation of a solid catalyst capable of being employed for the polymerization of the olefins especially ethylene, said catalyst comprising atoms of Mg, Cl, an electron donor compound such as tetrahydrofuran, a neutral metallocene and a solid support based on a refractory oxide.

Accordingly, the present invention provides a process for the preparation of supported metallocene catalyst which comprises, preparing a precursor catalyst material which consists of preparing the solution of magnesium halide compound in an electron donor solvent in which the magnesium halide will be completely soluble, heating the solution to a temperature in the range of 65°C to the boiling point of the respective electron donor for a period ranging between 10 to 60 minutes, separately preparing the solution of the metallocene compound into the same electron donor solvent, heating the solution to a temperature in the range of 25° C to 70° C for a period ranging between 0.1 to 0.5 hrs., mixing the two solutions to obtain the homogeneous solution of catalyst precursor compound, stirring and maintaining this resulting homogeneous solution at a temperature in the range of 50° C to 70° C for a period of 0.5 to 2 hrs., separately preparing the slurry of the support in the liquid hydrocarbon medium, heating the slurry to a temperature in the range of 50° C to 75° C and stirring it for a period of 0.25 to 2 hrs., mixing the homogeneous solution of the catalyst precursor compound with the slurry of the support, stirring the slurry for 0.5 to 3 hrs maintaining at a temperature ranging between 50° C to 70° C, cooling the resultant slurry to room temperature under inert atmosphere, separating the solid by conventional methods, washing the solid by hydrocarbon solvent, drying the solid under vacuum to obtain the product.

In an embodiment of the present invention, the electron donor solvent used for the preparation of the solution of the magnesium halide may be such as alkyl esters of aliphatic and aromatic carboxylic acids, aliphatic ethers, cyclic ethers and aliphatic ketones preferably alkyl esters of C1 to C4 saturated aliphatic carboxylic acids; alkyl esters of C7 to C8 aromatic carboxylic acids; C2 to C6, C3 to C4 aliphatic ethers; C3 to C4 cyclic ethers, C4 mono or di ether, most preferred being methyl formate, ethyl acetate, butyl acetate, hexyl ether, tetrahydrofuran, dioxane.

In another embodiment the magnesium halide used

may be such as magnesium chloride, magnesium bromide and magnesium iodide, preferably magnesium chloride.

In one another embodiment, the metallocene compound used has a general formula



Cp and Cp' designate each an unsaturated hydrocarbon radical with a central atom M. The groups Cp and Cp' can be obtained by a covalent bridge (bond).

- M indicates the transition metal which are chosen from the groups IIIB, IVB, VB and VIB of the Periodic Table.
- a, b and x designate the integral numbers such as $a+b+x = m$, $x > 0$, and a and / or b not equal to zero.
- m indicates the Valency of the transition metal M
- X designates a halogen selected from Cl, Br or I.

In yet another embodiment of the present invention, the transitional element in the metallocene compound may be such as scandium, titanium, zirconium, hafnium and vanadium, preferably zirconium.

In still another embodiment of the present invention, the Cp and Cp' groups in the metallocene compound are mono or a polycyclic group substituted with 5 to 50 carbon atoms bond with a double conjugated bond such as cyclopentadienyl, indenyl, or a fluorenyl radical or a derivative substituted by this radical containing upto 10 atoms of carbon, or a radical derived from the elements chosen from the group VA of the Periodic Table, such as N or P.

Refractory oxide contains hydroxyl functional groups and may have a specific surface area (BET) of 50 to 1000 m²/g, especially, from 100 to 400 m²/g and a pore vol. of 0.5 to 5 ml/g preferably, from 1 to 3 ml/g.

In yet another embodiment of the present invention, the support may be selected from inert porous materials such as dry powders of oxides of silicon or aluminum having an average particle size of about 10 to 250 microns, preferably 50 to 150 microns and a surface area of 50 to 600 m²/g, preferably 100 to 300 m²/g.

In one more embodiment of the present invention, the ratio of magnesium halide to per mole of the metallocene compound is 0.5 to 50 w/w preferably 1 to 8.

In another embodiment, the (Cp)_a(Cp')_bM X_x may be selected from Cp₂ZrC₁₂ and bisindenyl zirconium dichloride. The amount of magnesium halide onto the carrier may be in the range of 0.5 to 50 wt%.

In a feature of the present invention, the inert support materials should be dry i.e. free of absorbed water. Drying of the carrier materials can be effected by heating at a temperature of about 150°C or more than that or by treating these materials by means of aluminum alkyls compounds or aluminoxanes compounds. Higher the calcining temperature, lower is the amount of hydroxyl

group on the support. Temperature of about 150 to 900°C can be used with the preferred range between 400 to 600°C. Modification of the support in this manner provides the catalyst composition with increased activity.

The quantity of hydroxyl groups in the support depends on the support employed, as its specific area, on the physicochemical treatment and on drying to which it may have been subjected beforehand. A support which is ready for use generally contains from 0.1 to 5 preferably, from 0.5 to 3 mmol of hydroxyl groups per gram.

In a feature of the present invention, the support which may be granular, is preferably devoid of free water at the time of its use in the catalyst preparation. For this purpose, it can be preferably made free water by means which are well known, such as heat treatment ranging from 100 to 950 °C specially, 150 to 700 °C or by treating the oxide material with aluminum alkyl or aluminoxane compounds. The support may be chosen, in particular, from a silica, alumina, silica-alumina or a mixture of these oxides, and may consist of particles which have a mass mean diameter ranging from 30 to 240 microns especially, from 50 to 150 microns. The use of silica is preferred, specially ones sold by Davison Chemical under the commercial reference Davison 952.

In another feature of the present invention, the essential characteristics of a solid catalyst precursor is the halide state of neutral metallocene containing at least one cyclopentadienyl ligand in the backbone moiety making a part of the precursor.

Neutral metallocene with formula ((Cp)_a(Cp')_bMX_x) may be cited as an example, in which X is selected from Cl, Br or I. The mono and dihalide scandium metallocenes such as chlorodi(cyclopentadienyl)scandium, and dichloro(indenyl)scandium, mono, di and trihalides titanium metallocenes, such as chloro-tri(pentamethyl cyclopentadienyl)titanium, trichloro(cyclopentadienyl)titanium, mono, di or trihalides of zirconium such as trichloro(cyclopentadienyl)zirconium, dichloro(biscyclopentadienyl)zirconium metallocene. Among which, the last one is the most preferred.

Yet another feature of the present invention is that the electron donor compound should be a liquid organic compound at 25°C and in which the metallocene and the magnesium compound are partially or completely soluble.

In yet another feature of the present invention the precursor is formed by dissolving at least one metallocene compound and at least one magnesium compound in at least one electron donor compound at a temperature of from about 20°C upto the boiling point of the electron donor compound. The metallocene compound can be added to the electron donor compound before or after addition of the magnesium compound. The dissolution of the magnesium and the metallocene compound can be facilitated by stirring, or by refluxing the two compounds in the electron donor compound.

The solution of these two components is impregnated into the porous inert carrier material in a weight ratio of about 0.033:1 to 1:1 and preferably from 0.15:1 to 0.33:1 parts of Mg/Zr/electron donor composites per part by weight of carrier material.

In yet another feature of the present invention the impregnation of the dried support may be accompanied by addition of an aliphatic hydrocarbon solvent to the inert carrier material at a temperature of about 25°C to near boiling point of the solvent and then admixing the previous magnesium and metallocene compound dissolved in electron donor to the porous carrier material to impregnate the support. The catalyst may be washed several times with an aliphatic hydrocarbon. The excess aliphatic hydrocarbon is then removed by drying under vacuum until a constant weight is obtained. Drying can be done at a temperature from about 40°C to 80°C, especially, at 25°C to 35°C.

In yet another feature of the present invention, the catalyst precursor made according to the present invention has the formula



wherein,

X is selected from the group consisting of Cl, Br, I and may be mixtures thereof.

M indicates the transition metal which are chosen from the groups IIB, IVB, VB, VIB of the Period Table.

ED is an electron donor compound,

a is 1 to 50, preferably, 1.5 to 2.5

b is 0.01 to 1 preferably, 0.02 to 0.05

c is 2 to 80, preferably, 1.5 to 2.5

d is 0.03 to 1, preferably, 0.05 to 0.09

The element Mg is the Arabic numerical one and the amount of magnesium chloride on the carrier material is in the range preferably from 0.5 to 50, especially, from 1 to 5 weight percent.

These catalysts are useful for the polymerization and copolymerization of olefins as described in the co-pending Indian patent application no. which corresponds to European Patent Application No

The process of the present invention is described hereinbelow with examples

which are illustrative only and should not be construed to limit the scope of the present invention in any manner.

Example-1

This example illustrates the preparation of the catalyst precursor:

All glass equipment was heated under vacuum and flushed with nitrogen. All manipulations involving air-sensitive compounds were performed inside a Lab-

conco Model 50004 inert atmosphere glove box continuously purged with high purity N₂ from a generator (Spartech Model NG 300-1) or under a positive pressure of high purity N₂ using standard bench top inert atmosphere techniques. The solvent n-hexane and electron donor used in each case freshly distilled over sodium under N₂. Magnesium was estimated titrimetrically using EDTA. Chlorine was estimated by argentometric method. The amount of zirconium in the catalyst was determined by using Inductively Coupled Plasma ionization spectrometer.

Example-2

A support, namely, "Davison 952" silica powder sold by Davison Chemical Company which had a specific surface area of 300 m²/g and a pore volume 1.65 cc/g, hydroxyl groups, 2.8 mmol / g of silica was employed. It was dried at 550°C for 14 hours to obtain a silica powder free of water and containing 1 mmol hydroxyl groups per gram of silica.

In a three neck round bottom flask equipped with magnetic needle, reflux condenser, N₂ inlet and outlet which was flame dried under vacuum and cooled under nitrogen. Anhydrous MgCl₂, 394.3 mg was placed followed by addition of 70 ml of tetrahydrofuran. The resulting mixture was then heated at 60-65°C or 1/2 h until all magnesium chloride dissolved in tetrahydrofuran. Bis(cyclopentadienyl) zirconium (IV)dichloride, (335 mg) was dissolved in 20 ml of tetrahydrofuran and the solution was added to the solution of magnesium chloride over a period of 15 minutes. The resulting homogeneous solution was stirred for 1/2 h at 60-65°C.

In another three neck flask which was flame dried and cooled under N₂, 4.22 g of calcined silica was placed followed by 300 ml of n-hexane. The slurry was stirred for 1/2 h at 60-65°C.

The homogeneous solution of magnesium chloride and zirconium metallocene in tetrahydrofuran prepared as above was added to the slurry of silica solution over a period of 1/2 h. The whole slurry was stirred for 1 h at 60-65°C.

The slurry was cooled to room temperature under N₂, filtered and washed three times each with 50 ml of n-hexane. The solid was dried under vacuum. The amount of Zr, Cl, Mg and THF (wt %) in the resultant catalyst solid was 0.36, 10, 2.4 and 15.7 respectively.

Example-3

A support, namely, "Davison 952" silica powder sold by Davison Chemical Company which had a specific surface area of 300 m²/g and a pore volume 1.65 cc / g, hydroxyl groups, 2.8 mmol / g of silica was employed. It was dried at 550 °C for 14 hours to obtain a silica powder free of water and containing 1 mmol hydroxyl groups per gram of silica.

In a three neck round bottom flask equipped with

magnetic needle, reflux condenser, N₂ inlet and outlet which was flame dried under vacuum and cooled under nitrogen. Anhydrous MgCl₂, 209.5 mg was placed followed by addition of 40 ml of tetrahydrofuran. The resulting mixture was then heated at 60-65 °C for 1/2 h until all magnesium chloride dissolved in tetrahydrofuran. Bis(cyclopentadienyl) zirconium (IV)dichloride, (251-mg) was dissolved in 10 ml of tetrahydrofuran and the solution was added to the solution of magnesium chloride over a period of 25 minutes. The resulting homogeneous solution was stirred for 1/2 h at 60-65°C.

In another three neck flask which was flame dried and cooled under N₂, 2.54 g of calcined silica was placed followed by 200 ml of n-hexane. The slurry was stirred for 1/2 h at 60-65°C.

The homogeneous solution of magnesium chloride and zirconium metallocene in tetrahydrofuran prepared as above was added to the slurry of silica solution over a period of 1/2 h. The whole slurry was stirred for 1 h at 60-65°C.

The slurry was cooled to room temperature under N₂, filtered and washed three times each with 30 ml of n-hexane. The solid was dried under vacuum. The amount of Zr, Cl, Mg and THF (wt %) in the resultant catalyst solid was 0.45, 10.7, 2.0 and 17 respectively.

Example-4

A support, namely, "Davision 952" silica powder sold by Davision Chemical Company which had a specific surface area of 300 m²/g and a pore volume 1.65 cc / g, hydroxyl groups, 2.8 mmol / g of silica was employed. It was dried at 550 °C for 14 hours to obtain a silica powder free of water and containing 1 mmol hydroxyl groups per gram of silica.

In a three neck round bottom flask equipped with magnetic needle, reflux condenser, N₂ inlet and outlet which was flame dried under vacuum and cooled under nitrogen. Anhydrous MgCl₂, 180.3 mg was placed followed by addition of 40 ml of tetrahydrofuran. The resulting mixture was then heated at 70-75 °C for 1/2 h until all magnesium chloride dissolved in tetrahydrofuran. Bis(cyclopentadienyl) zirconium (IV)dichloride, (255 mg) was dissolved in 10 ml of tetrahydrofuran and the solution was added to the solution of magnesium chloride over a period of 20 minutes. The resulting homogeneous solution was stirred for 1/2h at 70- 75°C.

In another three neck flask which was flame dried and cooled under N₂, 2.50 g of calcined silica was placed followed by 200 ml of n-hexane. The slurry was stirred for 1/2 hat 70- 75°C.

The homogeneous solution of magnesium chloride and zirconium metallocene in tetrahydrofuran prepared as above was added to the slurry of silica solution over a period of 1/2 h. The whole slurry was stirred for 1 h at 70-75°C.

The slurry was cooled to room temperature under N₂, filtered and washed three times each with 30 ml of

n-hexane. The solid was dried under vacuum. The amount of Zr, Cl, Mg and THF (wt %) in the resultant catalyst solid was 0.48, 9.70, 1.8 and 15.0 respectively.

Example-5

A support, namely, "Davision 952" silica powder sold by Davision Chemical Company which had a specific surface area of 300 m²/g and a pore volume 1.65 cc/g, hydroxyl groups, 2.8 mmol / g of silica was employed. It was dried at 550°C for 14 hours to obtain a silica powder free of water and containing 1 mmol hydroxyl groups per gram of silica.

In a three neck round bottom flask equipped with magnetic needle, reflux condenser, N₂ inlet and outlet which was flame dried under vacuum and cooled under nitrogen. Anhydrous MgCl₂, 200 mg was placed followed by addition of 50 ml of tetrahydrofuran. The resulting mixture was then heated at 60-65°C for 1/2 h until all magnesium chloride dissolved in tetrahydrofuran. Bis(cyclopentadienyl) zirconium (IV)di chloride, (250 mg) was dissolved in 15 ml of tetrahydrofuran and the solution was added to the solution of magnesium chloride over a period of 15 minutes. The resulting homogeneous solution was stirred for 1/2 h at 60-65°C.

In another three neck flask which was flame dried and cooled under N₂, 2.60 g of calcined silica was placed followed by 300 ml of n-hexane. The slurry was stirred for 1/2 hat 60- 65°C.

The homogeneous solution of magnesium chloride and zirconium metallocene in tetrahydrofuran prepared as above was added to the slurry of silica solution over a period of 1/2 h. The whole slurry was stirred for 2 h at 60-65°C.

The slurry was cooled to room temperature under N₂, filtered and washed three times each with 50 ml of n-hexane. The solid was dried under vacuum. The amount of Zr, Cl, Mg and THF (wt %) in the resultant catalyst solid was 0.40, 10, 2.2 and 17 respectively.

Example-6

A support, namely, "Davision 952" silica powder sold by Davision Chemical Company which had a specific surface area of 300 m²/g and a pore volume 1.65 cc/g, hydroxyl groups, 2.8 mmol / g of silica was employed. It was dried at 550°C for 14 hours to obtain a silica powder free of water and containing 1 mmol hydroxyl groups per gram of silica.

In a three neck round bottom flask equipped with magnetic needle, reflux condenser, N₂ inlet and outlet which was flame dried under vacuum and cooled under nitrogen. Anhydrous MgCl₂, 277.6 mg was placed followed by addition of 40 ml of tetrahydrofuran. The resulting mixture was then heated at 60-65 °C for 1/2 h until all magnesium chloride dissolved in tetrahydrofuran. Bis(cyclopentadienyl) titanium (IV) dichloride, (503 mg) was dissolved in 20 ml of tetrahydrofuran and

th solution was added to th solution of magnesium chloride over a period of 15 minutes. The resulting homogeneous solution was stirred for 1/2 h at 60-65°C.

In another three neck flask which was flame dried and cooled under N₂, 2.90 g of calcined silica was placed followed by 300 ml of n-hexane. The slurry was stirred for 1/2 hat 60- 65°C.

The homogeneous solution of magnesium chloride and zirconium metallocene in tetrahydrofuran prepared as above was added to the slurry of silica solution over a period of 1/2 h. The whole slurry was stirred for 1 hat 60-65°C.

The slurry was cooled to room temperature under N₂, filtered and washed three times each with 50 ml of n-hexane. The solid was dried under vacuum. The amount of Ti, Cl, Mg and THF (wt %) in the resultant catalyst solid was 2.64 11.2, 2.3 and 14.2 respectively.

Example-7

A support, namely, "Davison 952" silica powder sold by Davison Chemical Company which had a specific surface area of 300 m²/g and a pore volume 1.65 cc/g, hydroxyl groups, 2.8 mmol / g of silica was employed. It was dried at 550°C for 14 hours to obtain a silica powder free of water and containing 1 mmol hydroxyl groups per gram of silica.

In a three neck round bottom flask equipped with magnetic needle, reflux condenser, N₂ inlet and outlet which was flame dried under vacuum and cooled under nitrogen. Anhydrous MgCl₂, 210 mg was placed followed by addition of 70 ml of tetrahydrofuran. The resulting mixture was then heated at 60-65°C for 1/2 h until all magnesium chloride dissolved in tetrahydrofuran. Bis(indenyl) zirconium (IV) dichloride, (300 mg) was dissolved in 20 ml of tetrahydrofuran and the solution was added to the solution of magnesium chloride over a period of 15 minutes. The resulting homogeneous solution was stirred for 1/2 hat 60- 65°C.

In another three neck flask which was flame dried and cooled under N₂, 2.4 g of calcined silica was placed followed by 300 ml of n-hexane. The slurry was stirred for 1/2 h at 60-65°C.

The homogeneous solution of magnesium chloride and zirconium metallocene in tetrahydrofuran prepared as above was added to the slurry of silica solution over a period of 1/2 h. The whole slurry was stirred for 2 h at 60-65°C.

The slurry was cooled to room temperature under N₂, filtered and washed three times each with 30 ml of n-hexane. The solid was dried under vacuum. The amount of Zr, Cl, Mg and THF (wt %) in the resultant catalyst solid was 0.30, 9.50, 2.4 and 16 respectively.

Claims

1. A process for th preparation of supported metallocene catalyst, which comprises preparing a pre-

cursor catalyst material which consists of preparing the solution of magnesium halide compound in an electron donor solvent in which the magnesium halide being completely soluble, heating the solution to a temperature in the range of 65°C to the boiling point of the respective electron donor for a period ranging between 10 to 60 minutes, separately preparing a solution of the metallocene compound into the same electron donor solvent, heating the solution to a temperature in the range of 25° C to 70° C for a period ranging between 0.1 to 0.5 hrs., mixing the two solutions to obtain the homogenous solution of catalyst precursor compound, stirring and maintaining this resulting homogenous solution at a temperature in the range of 50° C to 70° C for a period of 0.5 to 2 hrs., separately preparing the slurry of the support in the liquid hydrocarbon medium, heating the slurry to a temperature ill the range of 50° C to 75° C and stirring it for a period of 0.25 to 2 hrs., mixing the homogenous solution of the catalyst precursor compound with the slurry of the support, stirring the slurry for 0.5 to 3 hrs maintaining at a temperature ranging between 50° C to 70° C , cooling the resultant slurry to room temperature under inert atmosphere, separating the solid by conventional methods, washing the solid by hydrocarbon solvent, drying the solid under vacuum to obtain the product.

2. A process as claimed in claim 1 wherein the electron donor solvent used for the preparation of the solution of the magnesium halide is such as alkyl esters of aliphatic and aromatic carboxylic acids, aliphatic ethers, cyclic ethers and aliphatic ketones preferably alkyl esters of C1 to C4 saturated aliphatic carboxylic acids; alkyl esters of C7 to C8 aromatic carboxylic acids; C2 to C6, C3 to C4 aliphatic ethers; C3 to C4 cyclic ethers, C4 mono or di ether, most preferred being methyl formate, ethyl acetate, butyl acetate, hexyl ether, tetrahydrofuran, dioxane.
3. A process as claimed in claim 1 wherein the magnesium halide used is such as magnesium chloride, magnesium bromide and magnesium iodide, preferably magnesium chloride.
4. A process as claimed in claim 1 wherein the metallocene compound used has a general formula



where Cp and Cp' designate each an unsaturated hydrocarbonic radical with a central atom M. The groups Cp and Cp' can be obtained by a covalent bridge (bond).

- M indicates the transition metal which are cho-

sen from the groups IIIB, IVB, VB and VIB of the Periodic Table.

- a, b and x designate the integral numbers such as $a+b+x = m$, $x > 0$, and a and / or b not equal to zero.
- m indicates the Valency of the transition metal M
- X designates a halogen selected from Cl, Br or I.

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5. A process as claimed in claim 1 wherein the transitional element in the metallocene compound is such as scandium, titanium, zirconium, hafnium and vanadium, preferably zirconium.

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6. A process as claimed in claim 1 wherein the Cp and Cp' groups in the metallocene compound are mono or a polycyclic group substituted with 5 to 50 carbon atoms bond with a double conjugated bond such as cyclopentadienyl, indenyl, or a fluorenyl radical or a derivative substituted by this radical containing upto 10 atoms of carbon, or a radical derived from the elements chosen from the group VA of the Periodic Table, such as N or P.

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7. A process as claimed in claim 1 wherein the support is selected from inert porous materials such as dry powders of oxides of silicon or aluminum having an average particle size of about 10 to 250 microns.

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8. A process as claimed in claim 1 wherein the ratio of magnesium halide to per mole of the metallocene compound is 0.5 to 50 w/w.

9. A process as claimed in claim 1 wherein the $(Cp)_a(Cp')_bMX_x$ is selected from Cp_2ZrCl_2 , Cp_2TiCl_2 , and bisindenyl zirconium dichloride.

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10. A process as claimed in claim 1 wherein the amount of magnesium halide onto the carrier material is in the range of 0.5 to 50 weight percent.

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European Patent
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EUROPEAN SEARCH REPORT

Application Number
EP 97 89 0088

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int.Cl.6)
X	EP 0 416 928 A (NESTE OY) * claims; page 3, lines 16-20; page 3, line 52 to page 4, line 17 *	1-10	C08F4/654
A	EP 0 708 116 A (ATOCHEM ELF SA) * claims; page 3, lines 34-48; page 8, example 14 *	1-10	
D,A	EP 0 613 908 A (SOLVAY) * claims; page 12, example 1 *	1-10	
			TECHNICAL FIELDS SEARCHED (Int.Cl.6)
			C08F
The present search report has been drawn up for all claims			
Place of search THE HAGUE		Date of completion of the search 24 June 1997	Examiner Mergoni, M
<p>CATEGORY OF CITED DOCUMENTS</p> <p>X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure F : intermediate document</p> <p>T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons</p> <p>Δ : member of the same patent family, corresponding document</p>			

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